

In this expression, the magnitudes indicated by the capital letters may belong to the solvent, those indicated by small letters to the substance dissolved. The vapor-tension S would, of course, be that, not of the pure solvent but as diminished by the material in solution, according to Raoult's law.

In practice the method thus outlined may be carried out as follows: Assume, as a concrete case, that a determination of the tension of iodine in a solution of this element with potassium iodide and water is to be undertaken. The solution is weighed in a suitable vessel, and placed in a thermostat. A slow current of dry air is passed through it for a time, and it is again weighed. The loss of weight equals evaporated water plus iodine, (or $= w + W$), and titration of the solution before and after shows how much of this loss is due to evaporated iodine.

All the required data for formula (3) are therefore at hand. As yet, the method has been applied only to a study of starch iodide, of which an account will be found in the following paper. It was devised and experiments with it were begun more than two years ago, but circumstances beyond control of the present writer have delayed publication so long already that it seems undesirable to withhold preliminary notice any longer.

So far as iodine solutions are concerned, I would reserve the field for a time.

CONTRIBUTIONS TO THE STUDY OF STARCH IODIDE.

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SINCE the discovery of the blue substance known as "starch iodide" by Stromeyer,¹ our knowledge of its properties has been enriched by numerous researches². In spite, however, of the remarkably extensive literature pertaining to the subject, no consensus of opinion has been arrived at with regard to either the nature of the substance, its composition, or even some of its more important properties.

Thus, while Blondlot, Béchamp, Pohl, R. Fresenius, Duclaux, Brückner, and Küster² consider it to be a mixture of starch with

¹ Thorpe's Dict., Vol. III, p. 565.

² See bibliography at the end of this paper.

iodine, or a solution of the latter in the former, Rouvier, Payen, Fritsche, Bondonneau, P. Guichard, Sonstadt, Pellet and others hold it to be a chemical compound of starch and iodine. These authorities are nevertheless far from agreement as to the composition of the compound. They give the amount of iodine contained all the way from 3.2 per cent. as a minimum, up to 19.6 per cent. as a maximum, and present formulas varying over a correspondingly wide range.

Ostwald, in the first edition of his "Wissenschaftliche Grundlagen," holds it to be an easily dissociable compound, while in the second edition of the same book, influenced probably by Küster's recent work, he adds the modifying clause "beziehungsweise einer festen Lösung." Bernthsen¹ takes an identical view,² regarding starch iodide as a dissociable compound.

Mylius³ maintains that the presence of hydriodic acid or an iodide is essential to the formation of starch iodide, to which he assigns the formula $(C_{24}H_{40}O_{20}I)_4HI$ as most probable. Seyfert, as also Tóth, disputed the presence of hydriodic acid or alkali iodide as essential to the existence of the blue compound.

EXPERIMENTAL.

It seemed best to simplify the conditions so far as possible by confining the study of the action of iodine on starch to clear solutions of the latter, in order to avoid complications arising from variations in the physical condition of the starch. Therefore, in the work described in the following pages, only such solutions were used.

The iodine employed was prepared from "chemically pure" iodine of the shops by distillation with potassium iodide in a current of steam. It was free from cyanide.

The solutions of starch were made by heating maize starch⁴ with water in an autoclave. At first a few drops of chloroform were added to the solutions as a preservative, but, as this seemed to induce partial precipitation of starch after standing, this practice was discarded after the first set of experiments described below.

¹ "Organische Chemie," sixth edition, p. 307.

² "Sowohl die Stärkekörner, als der Stärkekleister werden von Jod intensiv blau, von Brom feuergelb gefärbt, indem lockere, additionelle Verbindungen entstehen."

³ See bibliography.

⁴ Washed with water, dilute hydrochloric acid, alcohol, and water.

Starch Iodide Prepared in the Cold.

Series 1.—Ten grams of starch were heated to 175° with 100 cc. of water for about three hours. The solution was brownish but perfectly clear, and filtered readily. A few drops of chloroform were added, and after standing a day, part of the starch had precipitated. The solution, somewhat diluted, was well shaken with fine asbestos, when it easily filtered clear.

To determine the strength of this solution, 5 cc. were evaporated to dryness at 100° and brought to constant weight by heating in an air-bath at 105°-110°. This operation demands care, in consequence of the extraordinarily hygroscopic character of the product. Five cc. of solution yielded 0.1834 gram of starch.

This standard solution was diluted to make solutions containing respectively, 10.000, 5.000, 2.500, 1.250, and 0.625 grams per liter. To each solution was added an amount of pulverized iodine equal to the amount of starch present, and the whole was shaken by a motor (usually for about six hours), allowed to stand over night, a sample withdrawn for titration, and the remainder again shaken, etc. In the following table the amounts, in grams per liter, of iodine dissolved are shown.

TABLE I.

Concentration of starch, g. per L.	10	5 ¹	5	2.5	1.25	0.625
Concentration of iodine, g. per L. :						
After 5 days	0.744	0.474	0.468	0.290	0.264	0.242
After 7 days	0.842	0.520	0.526	0.348	0.290	0.252
After 10 days	0.894	0.552	0.566	0.342	0.299	0.262
After 11 days	0.922	0.572	0.598	0.408	0.318	0.270
After 12 days	0.936	0.580	0.598	0.402	0.322	0.286
After 17 days	1.028	0.648	0.652	0.428	0.318

If the results are plotted as a curve with the respective starch concentrations as ordinates and the final iodine concentrations (solubilities) as abscissas, practically a straight line is obtained. This line cuts the *x* axis at a point corresponding to the iodine concentration of 0.24 gram per liter, about, that is to say, at the point corresponding very closely with the best determinations of the solubility of iodine in water. The temperature fluctuated between 22° and 25°, 20° being only once recorded.

¹ This solution was freed from chloroform by blowing air through it before adding iodine. The figures show that the presence of traces of chloroform do not influence the solubility of the iodine.

Allowing for the solubility of the iodine in the water, we find a fairly steady relation between iodine and starch. The mean value of this relation is about 7.63 per cent., or in other words, the "starch iodide" in the solution contains about 7.1 per cent. of iodine.

Since the starch solution used in this series of experiments gave some evidence of having been heated to too high a temperature (it gave with a trace of iodine a violet instead of clear blue), we determined upon another series, in which the solutions should be made at a lower temperature. It was found that the starch dissolved completely at 144° but precipitated again on cooling, but at 152° (5 atmospheres' pressure) the solution remained clear long enough for the object in view.

Series 2.—Fifteen grams of starch were heated for two and a half hours to 152° with 130 cc. of water. The resulting colorless solution filtered readily and did not show any tendency to precipitate on cooling. It was diluted without delay to 650 cc., of which 5 cc. were evaporated as before to determine the starch. 0.1077 gram of dry starch was obtained.

From this standard solution, six others were at once made, without waiting for the result of the determination, in order to avoid the possibility of separation of starch. These contained in 200 cc., respectively, 200, 150, 100, 75, 50, and 25 cc. of standard solution, or, 21.54, 16.16, 10.78, 8.08, 5.38, 2.70 grams of starch per liter.

These solutions were shaken six hours a day for ten days with excess of iodine at 21° to 26°. They then contained no undissolved starch or starch iodide, and were filtered to remove any finely divided iodine in suspension, and titrated with N/100 sodium thiosulphate. The data so obtained are recorded in the following table.

TABLE II.

<i>C</i> . Concentration of starch, grams per liter.....	21.54	16.16	10.78	8.08	5.38	2.70
<i>C</i> . Concentration of iodine dissolved, grams per liter.....	1.644	1.316	0.970	0.786	0.602	0.446
<i>C</i> — <i>C</i> . Concentration of iodine corrected for solubility in water	1.376	1.048	0.702	0.518	0.334	0.178
<i>P</i> . Ratio of starch to iodine in percentage.....	6.39	6.48	6.51	6.41	6.21	6.59

The data given in the table under rubrics I and II were reduced by the formula $100 \frac{C - C_w}{C_s} = P$, in which

C_s = the concentration of the starch (rubric I) ;

C_w = the iodine dissolved in the water alone, assumed as constant ;

C = total concentration of dissolved iodine (rubric II) ;

P = per cent. of iodine taken up by the starch.

The numerical values of C_w and of P as calculated from the observations themselves are, respectively, 0.268 and 6.43. The former is in excellent agreement with the known solubility of iodine in water. The temperature varied between 21° and 26° . The percentage of iodine to starch found, 6.43, corresponds fairly to that required by the formula $(C_6H_{10}O_5)_{12}I$, which is 6.53.

Series 3.—The preceding experiments having established that the amount of iodine taken up by the starch is constant and independent of the concentration of the starch solution so long as the water is saturated by iodine,¹ it was decided to make a series of determinations of the distribution of iodine between starch iodide solutions and an immiscible solvent (chloroform).²

The starch iodide solutions (being identical with those used in the experiments of Series 2) were shaken with 10 per cent. (by volume) of carefully purified chloroform in stoppered glass cylinders, until a condition of equilibrium was attained, and were then allowed to stand for several hours, usually over night, to permit the liquids to separate completely. Five cc. of the clear supernatant blue solution were titrated with N/200 thiosulphate and the rest of it transferred to a dry cylinder and mixed as before with 10 per cent. of its volume of chloroform.

In this way, six successive distribution determinations were made on each solution with constant concentration as to starch but diminishing iodine. The temperatures did not vary more than 1° from 21° throughout the series.

¹ Küster (*Ann. Chem.* (Liebig), **360**, 283) concludes that the amount of iodine absorbed by *solid* starch is wholly dependent on the concentration of the solution in contact with it.

² Compare Jakovkin's study of the dissociation of potassium triiodide in aqueous solution (*Ztschr. phys. Chem.*, **13**, 539 (1894)).

TABLE III.

Distribution of Iodine between Starch Solution and Chloroform.

1.	2.	3.	4.	5.	6.
Extraction mm- ber.	Concentration of iodine in starch solution.	Concentration of iodine in chloro- form.	Concentration of starch.	Distribution coeffi- cient between so- lutions.	Distribution of iodine chloroform to starch.
	A.	B.	C.	B + A.	BC + A.
0	1.644	0.00	21.54
1	1.218	4.26	21.54	3.50	75.5
2	1.076	1.42	21.54	1.32	28.4
3	0.990	0.86	21.54	0.87	18.7
4	0.918	0.72	21.54	0.78	...
5	0.890	0.28	21.54	0.31	...
6	0.858	0.32	21.54	0.37	...
0	1.316	0.00	16.16
1	0.910	4.06	16.16	4.46	72.1
2	0.780	1.30	16.16	1.67	27.0
3	0.732	0.48	16.16	0.66	10.5
4	0.682	0.50	16.16	0.73	...
5	0.644	0.38	16.16	0.59	...
6	0.620	0.24	16.16	0.39	...
0	0.970	0.00	10.78
1	0.616	3.54	10.78	5.74	61.6
2	0.524	0.92	10.78	1.76	19.0
3	0.484	0.40	10.78	0.83	8.9
4	0.442	0.42	10.78	0.95	...
5	0.408	0.34	10.78	0.83	...
6	0.380	0.28	10.78	0.74	...
0	0.786	0.00	8.08
1	0.472	3.14	8.08	6.65	53.7
2	0.394	0.78	8.08	1.98	16.0
3	0.348	0.46	8.08	1.32	10.7
4	0.318	0.30	8.08	0.94	...
5	0.290	0.28	8.08	0.97	...
6	0.276	0.14	8.08	0.50	...
0	0.608	0.00	5.38
1	0.328	2.74	5.38	8.35	45.0
2	0.250	0.78	5.38	3.12	16.8
3	0.210	0.40	5.38	1.90	10.2
4	0.190	0.20	5.38	1.05	...
5	0.176	0.14	5.38	0.80	...
6	0.158	0.18	5.38	1.14	...

1. Extraction num- ber.	2. Concentration of iodine in starch solution.	3. Concentration of iodine in chloro- form.	4. Concentration of starch.	5. Distribution coeffi- cient between so- lutions. <i>B + A.</i>	6. Distribution of iodine, chloroform to starch. <i>BC + A.</i>
0	0.446	0.00	2.70
1	0.176	2.70	2.70	15.34	
2	0.122	0.54	2.70	4.42	41.5
3	0.098	0.24	2.70	2.45	11.9
4	0.86	0.12	2.70	1.40	5.4
5	0.72	0.14	2.70	1.94	...
6	0.66	0.06	2.70	0.91	...

All concentrations are given in grams per liter.

Inasmuch as the iodine in the chloroform was determined by difference, there is an accumulation of errors in the values found for the chloroform concentrations (column 3) for the last three extractions and in the corresponding distribution coefficients (column 5). It will be observed that there is no indication of a constant distribution coefficient between chloroform and starch solution or between chloroform and dissolved starch. The results show a more rapid removal of the iodine from the solutions containing less starch, in such a way that all the curves cross at nearly a common point, corresponding to about 5 per cent., falling between the first and second extractions.

Series 4. (Starch Iodide Prepared by Heating).—43.152 grams of air-dried starch, containing 38.449 grams of dry starch (maize), were introduced into a Jena glass tube with 7 grams of pure iodine and enough water to moisten the mass completely. The tube was then sealed and heated to 99° for two hours. At the end of this time the contents formed a soft mass which readily dissolved in water, leaving an insoluble residue consisting of the excess of iodine only, but containing no starch. The solution, diluted to 2 liters, was found on titration with N/100 thiosulphate, to contain 2.562 mg. "free" iodine per cubic centimeter. From this standard solution were prepared, by appropriate dilution, five others, the strength of each of which, expressed in grams per liter and also in molecular volume, is shown in the following table. In calculating the latter the molecular weight of

starch is arbitrarily assumed as that of the molecule $(C_6H_{10}O_5)_6 = 972.5$.

TABLE IV.

Solution No.	Starch, grams per liter.	Iodine, grams per liter.	Molecular volume of starch.	Molecular volume of iodine.
1	19.225	2.562	50.69	99.02
2	14.42	1.922	67.58	132.03
3	9.61	1.281	101.38	198.04
4	7.21	0.961	135.16	264.06
5	4.806	0.641	202.76	396.08
6	2.403	0.320	405.52	792.16

It is obvious that these solutions prepared by the action of iodine on starch at about 100° are entirely different in composition from those previously described, which were made at about 20° from dissolved starch. The former correspond very closely with the formula $(C_6H_{10}O_5)_{12} \cdot I_2$, whereas we have seen that the latter possesses a composition nearly agreeing with the formula $(C_6H_{10}O_5)_{12} \cdot I$.

The solutions were shaken with chloroform repeatedly, exactly as described on a previous page. Unfortunately the temperatures were not kept constant during the extractions, and the figures show that the results were influenced by this circumstance to a much greater degree than was anticipated. It is nevertheless thought worth while to record the data obtained, since they demonstrate that in spite of the difference in percentage of iodine in the starch iodide as prepared cold and hot respectively, yet the behavior of both kinds toward an immiscible solvent is similar.

TABLE V.

1. Extraction num- ber.	2. Iodine in starch solution.	3. Iodine in chloro- form.	4. Concentration of starch - C.	5. Distribution coef- ficient between solutions.	6. BC + A.	7. Temperature.
0	2.562	19.22
1	1.960	6.016	19.22	3.06	58.9	23°
2	1.816	1.450	19.22	0.80	15.4	21°
3	1.792	0.234	19.22	0.13	2.5	18°
4	1.740	0.514	19.22	0.29	5.6	19°
5	1.684	0.562	19.22	0.33	6.5	18.5°

1. Extraction num- ber.	2. A. Iodine in starch solution.	3. B. Iodine in chloro- form.	4. C. Concentration num- of starch — C.	5. B + A. Distribution coef- ficient between solutions.	6. BC + A.	7. Temperature.
0	1.922	14.42
1	1.414	5.068	14.42	3.58	51.6	25°
2	1.308	1.068	14.42	0.82	11.7	21°
3	1.290	0.172	14.42	0.13	1.9	18°
4	1.244	0.460	14.42	0.37	5.4	19°
5	1.212	0.328	14.42	0.27	3.9	18.5°
0	1.281	9.61
1	0.904	3.782	9.61	4.45	42.8	25°
2	0.824	0.796	9.61	0.96	9.2	21°
3	0.804	0.196	9.61	0.24	2.3	18°
4	0.778	0.262	9.61	0.34	3.3	19°
5	0.756	0.220	9.61	0.29	2.8	18.5°
0	0.961	7.206
1	0.650	3.102	7.206	4.77	34.4	25°
2	0.594	0.562	7.206	0.95	6.7	21°
3	0.582	0.118	7.206	0.20	1.4	18°
4	0.554	0.284	7.206	0.51	3.7	19°
5	0.546	0.090	7.206	0.16	1.2	18.5°
0	0.641	4.805
1	0.416	2.240	4.805	5.38	25.9	25°
2	0.370	0.460	4.805	1.24	6.0	21°
3	0.358	0.120	4.805	0.33	1.6	18°
4	0.336	0.222	4.805	0.66	3.2	19°
5	0.322	0.142	4.805	0.44	2.1	18.5°
0	0.320	2.403
1	0.188	1.32	2.403	7.02	16.8	25°
2	0.160	0.286	2.403	1.79	4.3	21°
3	0.152	0.078	2.403	0.51	1.2	18°
4	0.136	0.152	2.403	1.12	2.7	19°
5	0.132	0.050	2.403	0.38	0.9	18.5°

Series 5.—More significant than the last experiments and freer from sources of error, are a number of determinations of the vapor-tension of the iodine in solutions of starch iodide, carried out after the method described by one of us in the preceding paper.

A solution was used for this purpose which contained 19.224 grams of dry starch and 1.485 grams of "free" iodine per liter, and which was prepared after the method already described by

heating the constituents together to 99°. The liquid was placed in a gas-washing apparatus, immersed in a tank of water, and a slow current of dried air was aspirated through it and then through a second similar apparatus identically charged. The object of the second vessel is to serve as control, its constancy of weight demonstrating that the current of air was slow enough to be fully saturated with the vapors of both water and iodine at the prevailing tension. In practice, small losses in the weight of this vessel were disregarded, a special experiment having shown that the ratio between iodine and water evaporated is the same when the air is *nearly* (95 per cent.) saturated as when it is fully so.

The total amount of water plus iodine evaporated during each period is given by the loss in weight of the bulb. The amount of iodine evaporated is determined by titration of the solution by N/100 thiosulphate at the close of each period of evaporation, and the water estimated by difference. The data of one such series of measurements are reproduced in Table VI. The tension in millimeters (s) of the iodine is found by the formula

$$s = 0.071 \times \frac{wS}{W} \text{ (see preceding paper), in which}$$

S = the tension of water-vapor at the prevailing temperature, and

w = the weight of water evaporated, and

W = the weight of iodine evaporated.

To secure greater accuracy, for each titration about 5 cc. of the solution were withdrawn and its weight determined. The iodine concentrations are accordingly given in milligrams per gram.

The minuteness of the tension of the iodine (less than 0.01 mm.) in solutions where the iodine forms 5 per cent. of the weight of the starch present is remarkable. To obtain correspondingly low tensions of iodine in solutions of potassium iodide, the amount of iodine must be reduced relatively to a much greater degree, but the general character of the curve is very similar. We have not as yet been able to calculate any satisfactory dissociation constant from the data. All the results obtained, both by the vapor-tension method and by the chloroform method, show the complexity of the phenomena involved, and lead to the hypothesis that the iodine in "starch iodide" solutions exists in two forms of union at least, a small portion of it being much more readily removed than the rest.

TABLE VI.

Table Showing Relative Evaporation of Water and of Iodine from Solutions of "Starch Iodide."

Period.....	1	2	3	4	5	6	7	8	9
Date	May 16	May 16	May 17	May 17	May 17	17-18	18	18-19	19-21
Time in minutes	5	10	30	150	150	960	420	1290	1800
Mean temperature	22.66°	22.63°	21.89°	21.79°	21.69°	21.00°	20.30°	19.46°	18.40°
Initial weight of solution	179.058	164.960	159.805	154.435	149.147	143.930	138.151	132.701	126.416
Final weight of solution.....	179.050	164.905	159.532	154.318	149.023	143.243	137.777	131.546	124.903
Loss in grams by evaporation = $w + W$...	0.008	0.055	0.273	0.117	0.124	0.687	0.374	1.155	1.513
Mean concentration in milligrams of iodine per gram	1.4427	1.4190	1.3331	1.2418	1.2042	1.1303	1.0486	1.0027	0.9566
Iodine evaporated in milligrams = w	2.25	5.73	22.30	7.10	4.76	17.58	6.70	7.47	7.07
Water evaporated in milligrams = W	5.75	49.3	251	110	119	669	367	1148	1506
Tension of water-vapor = S	20.46	20.43	19.53	19.41	19.29	18.50	17.72	16.82	15.75
Tension of iodine-vapor = s	0.57	0.168	0.123	0.089	0.055	0.035	0.023	0.0078	0.0053

$$s = 0.071 \times \frac{wS}{W} = \text{tension in millimeters.}$$

The differences between the composition of the starch iodide prepared cold and that made by the aid of heat, invite a further examination of the action of high temperatures on this substance. The authorities disagree radically as to the phenomena resulting from the action of heat on solutions of this kind, some maintaining while others deny the permanent decolorization by heat when volatilization of iodine is prevented. We accordingly undertook an investigation of this matter.

Series 6.—A somewhat dilute solution of starch iodide sealed up in a tube of Jena glass and heated to 100° becomes colorless if very dilute, or straw color if less so. The color returns in its original intensity (colorimetric determination) if the tube is small and is rapidly heated to the required temperature and quickly cooled. But if the heating is continued, the color returns on cooling with ever feebler intensity, and at last fails to return at all. Seventy cc. of a solution of starch iodide (prepared like those referred to in Table V) were sealed in a Jena glass tube and heated in a steam-bath for several days, eight to ten hours each day. The solution did not lose its blue color at first (being too concentrated to do so) but the upper part of the tube showed the violet vapor of iodine. After four days the violet vapor had disappeared and the solution had become green. After heating fifty hours, the solution had become light yellow and the operation was discontinued. No pressure was apparent on opening the tube.

The solution was filtered from an insignificant greenish sediment. It was pale yellow, and gave the following reactions: Shaken with chloroform the latter becomes violet and the solution colorless. Starch paste turns it blue. Hence, free iodine is present and starch¹ is absent. Iodine dissolved very copiously in it, giving a dark brown solution. Methyl orange indicated a strongly acid reaction. Potassium nitrite set free a considerable quantity of iodine; hence, hydriodic acid is present. Mercurous nitrate gave a greenish precipitate of mercurous iodide. The solution reduced Fehling's solution, and on boiling with phenylhydrazine and acetic acid gave the crystalline precipitate characteristic of glucose.

For the purpose of quantitative study of these products, another

¹ Consequently the statement of Stocks (*Chem. News*, **57**, 183), that starch is unaltered on heating with iodine in a sealed tube, is incorrect.

portion of the material was sealed up and heated as before. The changes of color observed are noted below.

After heating five hours the color while hot was dark blue.

After heating fifteen hours the color while hot was black.

After heating twenty-five hours the color while hot was very dark green.

After heating thirty-five hours the color while hot was paler green.

After heating forty-five hours the color while hot was apple-green.

After heating fifty-five hours the color while hot and cold was yellow.

After heating sixty-five hours the color while hot and cold was colorless.

The product gave the same reactions as those described for the portion heated fifty hours, excepting the reactions of free iodine. Five cc. of the solution, titrated with N/100 potassium hydroxide and methyl orange, required 1.22 cc., which corresponds to 1.55 mg. of hydriodic acid. As the solution had previously been ascertained, by titration with thiosulphate, to contain in 5 cc. 12.81 mg. of iodine, it follows that 12.12 per cent. of the iodine originally present had, by continued heating, been converted into hydriodic acid.

Another portion of 5 cc. was neutralized with calcium carbonate and titrated with standard silver solution, using potassium chromate as indicator. We found thus in the 5 cc., 13.31 mg. of iodine. It follows that the greater part of the iodine is present in the form of a colorless organic iodide which gives up the whole of its iodine readily in the cold to a silver nitrate solution. The discrepancy between the 13.31 mg. found by the silver solution and the 12.81 mg. found by thiosulphate, is not due to experimental error, but is explained by the fact that solutions of starch iodide which have been made by heating for the normal period of about two hours contain a certain amount of the colorless organic iodide—an amount which one of us has found to average between three and four per cent. of the total iodine present. The organic iodide can not be removed even in traces from the solution which has been heated sixty-five hours, by shaking it with chloroform subsequent to neutralization. The existence of a colorless organic

iodide has been assumed, but not demonstrated, by Bondonneau,¹ by Duroy,² and by Guichard,³ although the last mentioned certainly came very near a demonstration by showing that starch iodide solutions, which have lost all color by evaporation of iodine, may be turned blue again by adding nitric acid, even when the hydriodic acid has (presumably) been removed by desiccation.

An important factor in the decolorization of starch iodide by prolonged heating is doubtless to be found in the saccharification of the starch by the hydriodic acid formed. That this is, however, not an essential factor was demonstrated by heating starch iodide with a small quantity of calcium carbonate in a glass tube. The heating had to be continued 110 hours to produce complete decolorization, whereas without the carbonate, under otherwise identical conditions, sixty-five hours sufficed. If starch iodide solutions are boiled in the open air until fully colorless, the resulting solution contains starch, hydriodic acid, colorless organic iodide, and glucose. Its color may be restored by addition of iodine or of iodate or of nitrous acid, etc.

While we are conscious that the question as to whether the substance called starch iodide is a compound or a solution of iodine in starch is still unsettled, we still believe that the observations recorded in this paper are more reconcilable with the former view than with the latter. If the blue color is due to the formation of a mere solution of iodine in starch, it is difficult to understand why iodine can not color dry starch blue, but the fact is that perfectly dry iodine imparts to perfectly dry starch a brownish color which becomes blue on contact with a trace of water. The presence of an organic solvent does not alter this phenomenon.

If the blue substance formed in presence of water is a mere solution of iodine in starch, what, we would like to ask, is the brown substance formed when iodine dissolves in starch in absence of water?

Potassium iodide, as is well known, in large excess imparts a brownish violet color to starch iodide solutions, dilution with water restoring the blue. As Mylius has observed, iodic acid changes the blue of starch iodide to brown, a fact which he uses to support the view that iodides must be present as an essential

¹ *Bull. Soc. Chim.*, **28**, 452 (1877).

² *Compt. rend.*, **51**, 1031 (1860); see also *Chem. News*, **3**, 16.

³ *Bull. Soc. Chim.*, **5**, 115 (1863).

part of the starch-blue. On this theory the change of color is simply caused by the removal of all iodide by the iodic acid. We have, however, found that if to the brown solution so formed, sodium or potassium bicarbonate or other weakly alkaline substances be added, the blue is immediately restored. This is inexplicable on the theory alluded to, but is easily understood if the possibility of iodine acting as a base is borne in mind, a possibility which has of late been shown by other authors¹ to be a probability. This involves the supposition that the iodic acid removes the iodine from the starch iodide in the form of iodine iodate,² and that when a stronger base is added, the weaker one (iodine) is displaced and placed at the disposition of the starch once more.

SUMMARY.

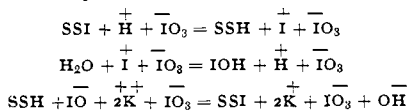
Clear starch solutions made at a temperature of about 150° take up in the cold an amount of iodine corresponding to the formula (C₆H₁₀O₅)₁₂I. Starch heated with excess of iodine to 100° for a short time takes up an amount of iodine corresponding to the formula (C₆H₁₀O₅)₁₂I₂. Starch and iodine heated to 100° for a longer time give ultimately a colorless solution which contains most of the iodine in the form of an organic iodide, a smaller part of it in the form of hydriodic acid, and which also contains glucose but no starch.

Solutions of starch iodide on shaking with chloroform yield, at first, considerable amounts of iodine to the latter but after removal of the first portion, give up iodine very slowly, showing no indication of a distribution coefficient of iodine between starch and chloroform as might be expected if the iodine were merely dissolved in the starch. The vapor-tension of iodine in solutions of starch iodide was determined by a new method, and was found to be extremely minute after the removal of the first portion of the iodine.

It is held that the substance known as starch iodide is to be

¹ Compare also for a similar assumption applied to chlorine, Noyes and Lyon : This Journal, 23, 463.

² The process imagined may be represented by ionic equations as follows, the symbol SS standing for the starch complex. :



looked upon as a dissociable additive compound of iodine with starch molecules, more or less depolymerized, according to the temperature employed, and subject therefore to variation in the ratio of starch to iodine. These views are supported by the following considerations:

1. Starch and iodine can not unite to produce the blue substance in the absence of water, a phenomenon frequently characteristic of chemical combination but not of solution in organic solvents.

2. The temporary decolorization of the blue solution by heat when loss of iodine is prevented is much better explained as a phenomenon of dissociation than of solution.

3. The fact that in the system, starch, water, chloroform, a rise of temperature has a very powerful effect in causing iodine to pass from the starch to the chloroform, indicates that the union of starch and iodine is exothermic, and agrees well with the hypothesis of a dissociable compound.

4. The fact that the amount of iodine taken up by the starch under different conditions, as to the temperature to which the starch is subjected either before or after iodine is added is naturally explained by the view that the molecule undergoes gradual depolymerization by steps, giving rise to two, or a series, of starch iodides of different composition. This also explains why no dissociation constant can be deduced.

5. The remarkably low vapor tensions of the iodine in these solutions, while quite natural if the iodine is combined, are exceptional and difficult to understand on any other basis.

The experimental part of this investigation was completed as here presented, two years ago, but publication was delayed by the temporary accidental loss of the records.

The authorities consulted (part of them it is proper to state by abstracts) are given in the following list, which is as complete a bibliography of the literature of starch iodide as we have been able to compile with the library facilities available.

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